

Name-

Date: 05-05-16

M.Marks: 20

Time- 40 min

Two Marks Each

Chemistry test

Class 10th ICSE



1. Explain with reasons the trends in metallic and non-metallic character down a group.
2. What are 'groups' of the Modern Periodic Table. What does the 'group number' signify.
3. State the relation between atomic number and atomic mass for light elements. State which elements are considered radioactive giving reasons.
4. Explain the trends from metallic to non-metallic character of the different elements in the first three periods.
5. Explain the trend in general of ionization potential of elements (a) on moving from left to right across a period (b) on moving down a group. Give reasons for the change in the periodic trend in each case.
6. What are 'periods'. State the correlation of a period number with the elements of that period.
7. Give reasons : Ionisation potential increases with increase in nuclear charge of the elements.
8. Give reasons : Atomic size decreases across a period but increases down a group of the periodic table.
9. Give reasons : Atomic size of group 18 [0 group] elements is more than the atomic size of group 17 [VIIA] elements.
10. Give reasons : Sulphur is placed in group 16 [VIA], chlorine in group 17 [VIIA] but argon in group 18 [0 group] of the periodic table.